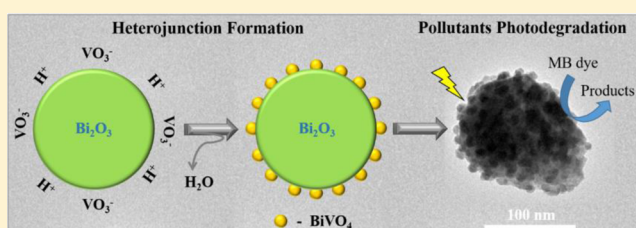


Growth of BiVO₄ Nanoparticles on a Bi₂O₃ Surface: Effect of Heterojunction Formation on Visible Irradiation-Driven Catalytic Performance

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S Supporting Information

ABSTRACT: Heterostructured materials composed of different semiconductors can be used to decrease rapid charge carrier recombination in photocatalysts, but the development of efficient synthesis methods for these materials remains a challenge. This work describes a novel strategy for tailoring heterostructures that is based on the solubility difference between two semiconductors with at least one metal in common. The growth of BiVO₄ on a preformed Bi₂O₃ particle was used as a model for heterojunction formation. The number of Bi₂O₃/BiVO₄ heterojunctions was tuned using synthesis variables (temperature and V concentration) and the particle size of the preformed Bi₂O₃. The synthesis of the Bi₂O₃/BiVO₄ heterostructures using Bi₂O₃ nanoparticles resulted in a larger quantity of heterojunctions due to the higher solubility of the nanoparticles compared to micrometric Bi₂O₃, which led to a classical heterogeneous precipitation on the preformed surfaces. The proposed growth mechanism was effective for obtaining heterostructured Bi₂O₃/BiVO₄ semiconductors with enhanced photocatalytic performances compared to the isolated phases. The greater photoactivity of the heterostructures could be explained by the increased spatial separation in the photogenerated electron/hole pairs due to the formation of a type-II heterostructure and was observed by time-resolved photoluminescence analysis. In this case, the photogenerated electrons were transferred from the conduction band of the p-type semiconductor (Bi₂O₃) to the n-type (BiVO₄) semiconductor, while the photogenerated holes were transferred from the valence band of the n-type semiconductor to the p-type semiconductor.



1. INTRODUCTION

The formation of heterostructures between semiconductors has been widely studied for applications in heterogeneous photocatalysis, such as organic pollutant degradation,^{1–6} water splitting,^{7–9} and artificial photosynthesis.^{10–13} This topic is of great interest in physical chemistry because it involves the potential to solve several problems related to water treatment and renewable energy, and the complete understanding of the charge transfer mechanism in heterostructures is required. A special feature of a suitable heterostructure is its capacity to increase the lifetime of a photogenerated electron/hole pair by suppressing its recombination.^{14–16} As a result, the charges instead migrate to the semiconductor surface and increase the occurrence of redox reactions on the heterostructure surface.^{17–24} In addition, the system formed between the BiVO₄ (type n)^{25,26} and Bi₂O₃ (type p) semiconductors^{27,28} exhibits promising electronic properties for the creation of type-II heterostructures (with p–n junctions).^{29,30} Due to its band gap characteristics, this heterostructure can be activated by visible

radiation, which is an important feature for photocatalytic applications under natural sunlight.^{25,26,31}

Heterostructured Bi₂O₃/BiVO₄ has been studied due to its potential for use in the photocatalytic degradation of organic pollutants^{32,33} and water splitting in photoelectrochemical (PEC) cells.³⁴ However, the formation of efficient heterojunctions between these two semiconductors remains a challenge.^{35,36} Recently, Ye et al.³⁴ proposed a synthesis method for building heterostructured Bi₂O₃/BiVO₄ films with a remarkable photocurrent in PEC cells. However, this method is based on multiple steps and favors the formation of a spurious phase (V₂O₅). Cheng et al.³² reported an interesting method for obtaining BiVO₄/Bi₂O₃ composites using a mild, one-step hydrothermal process, but the ratio between Bi₂O₃ and BiVO₄ could not be controlled. Other authors have used some additives/surfactants or a solvothermal method to obtain

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$\text{BiVO}_4/\text{Bi}_2\text{O}_3$. However, the use of a surfactant or solvent can poison the surface of the semiconductor.^{33,37,38} Additionally, the ratio of Bi_2O_3 to BiVO_4 and the morphology for either material was not controlled.^{32,33,37,38} In the work of Li and Yan,³⁹ $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures were obtained using surfactants with a calcination step at 600 °C, but the thermal treatment at high temperatures can negatively affect the properties of the nanocrystals due to surface dehydroxylation and surface area reduction. Therefore, the main challenge to overcome is the difficulty in controlling the simultaneous crystallization of two different components with suitable physical, chemical, and electronic properties through a surfactant-free and mild (low temperatures) synthesis method.^{17,36}

The use of a preformed particle to build heterostructures is of interest because the morphology and creation of interfaces can be easily controlled.^{17,36,40} Since Bi_2O_3 and BiVO_4 have Bi in common, it is possible to create an interface between the semiconductors by growing one phase on the sacrificial surface of a preformed particle using the difference in the solubility of the compounds. Therefore, the aim of this study was to develop a novel and efficient synthesis method to obtain $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures using a hydrothermal treatment and to evaluate the influence of the heterojunction on photocatalytic performance. The effects of the particle size (micrometric or nanometric) of the preformed Bi_2O_3 , the hydrothermal treatment temperature, and the amount of BiVO_4 grown on the Bi_2O_3 surface were investigated. The photoactivity of the heterostructures was probed using the photodegradation of methylene blue dye (MB) under visible irradiation. A mechanism was proposed for the effect of the charge transfer on the increased charge carrier lifetime of the type-II $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructure that was formed.

2. EXPERIMENTAL SECTION

Synthesis of $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ Heterostructures. The synthesis of the heterostructured $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ samples was performed by dispersing 0.2 g of either micrometric (Sigma-Vetec, 98%) or nanometric (Sigma-Aldrich, 90–210 nm particle size, 99.8%) preformed Bi_2O_3 precursor in 30 mL of distilled water, followed by the addition of an NH_4VO_3 precursor (water solubility at 25 °C is 7.81 g·L⁻¹) at different Bi:V molar ratios (1:1 or 1:2). The resulting reaction mixture was hydrothermally treated at 150 or 200 °C for 12 h. The materials obtained were washed with distilled water, centrifuged three times to remove impurities, and dried in an oven at 50 °C. Figure 1 shows a proposed mechanism for the growth of the BiVO_4 particles on the Bi_2O_3 sacrificial surface, which was driven by the solubility difference between the compounds. For

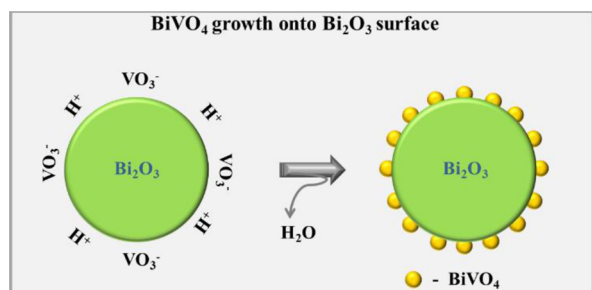


Figure 1. Proposed growth of BiVO_4 on a Bi_2O_3 surface.

comparative purposes, pure BiVO_4 was synthesized using a similar method with $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ and NH_4VO_3 in 30 mL of distilled water at a Bi:V molar ratio of 1:1, and this mixture was hydrothermally treated at 150 °C for 12 h.

The heterostructured samples are referred to as $s\text{-Het-}x:y T$, where s is the size scale of the Bi precursor (m for micrometric and n for nanometric), $x:y$ is the Bi:V molar ratio, and T is the crystallization temperature.

Powder Characterization. X-ray diffraction (XRD) patterns were recorded using a Shimadzu XRD 6000 diffractometer with Ni-filtered $\text{Cu K}\alpha$ ($\lambda = 0.15406$ nm) radiation. This instrument was operated at 30 kV and 30 mA in a continuous scanning mode at a speed of 2°·min⁻¹ with a step width of 0.02° from 10° to 60° 2θ . Raman spectroscopy measurements were performed using an FT-Raman spectrometer (Bruker RAM II with a Ge detector) equipped with a Nd:YAG laser with the wavelength centered at 1064 nm. UV–vis diffuse reflectance spectra (DRS) were recorded from 200 to 800 nm using a UV–vis spectrophotometer (Shimadzu UV-2600) equipped with an integrating sphere (ISR-2600 Plus) to determine the band gap of the materials. The measurements were performed in a total reflection mode using barium sulfate (BaSO_4) as the standard compound. The specific surface area (SSA) of the samples was calculated according to the BET model using the N_2 adsorption data obtained at -196 °C (Micrometrics ASAP 2000 instrument). All the samples were pretreated (degassed) via heating at 80 °C under vacuum until they reached a degassing pressure of less than 20 μmHg .

The morphology and size of the particles were investigated using field emission gun scanning electron microscopy (FE-SEM) (JEOL JSM 6701F). The semiquantitative atomic composition analysis and elemental mapping of Bi and V atoms were performed via energy-dispersive X-ray spectroscopy (EDX) using a Thermo Noran device coupled to a scanning electron microscope (JEOL JEM 2010). The formation of the heterostructures was confirmed via high resolution transmission electron microscopy (HRTEM) using a TECNAI G2 F20–LaB6 instrument operated at 200 kV. The samples were prepared for TEM by wetting carbon-coated copper grids with a drop of the colloidal alcoholic suspensions and then drying the samples in air.

Evaluation of the Photocatalytic Performance and Photodegradation Mechanism. The photocatalytic activities of the as-synthesized samples were evaluated using the photodegradation of methylene blue (MB) dye. In a typical experiment, 10 mg of the photocatalyst was added to 20 mL of a 5 mg·L⁻¹ MB dye aqueous solution. The dispersions were stirred and exposed to visible irradiation using six lamps (Osram, 15 W, maximum intensity at 440 nm) in a homemade photoreactor maintained at 18 °C. The photodegradation of MB was monitored at regular time intervals using UV–vis spectrophotometry (1601PC, Shimadzu) at an absorbance maximum of 665 nm. Prior to irradiation, the suspensions were maintained in the dark for 12 h with magnetic stirring to guarantee the complete adsorption–desorption equilibrium. The adsorption capacity of the MB dye for the samples was negligible (<5%) after 12 h.

The lifetimes of the charge carriers of the as-synthesized samples were determined via time-resolved photoluminescence using time-correlated single photon counting (TCSPC). A 405 nm pulsed laser diode (LDH P-C-405, PicoQuant) with an approximate 50 ps pulse width and 40 MHz repetition rate was used as the excitation source. The PL emission was spectrally

resolved using collection optics and an emission monochromator. The TCSPC module (PicoHarp 300, PicoQuant) was used for ultrafast detection. The deconvolution of the PL decay was performed using fitting software (FluoFit, PicoQuant) to deduce the time constant associated with the exponential decay.

3. RESULTS AND DISCUSSION

XRD patterns of the as-synthesized samples (Figure 2) were obtained to confirm the presence of both the BiVO_4 and Bi_2O_3

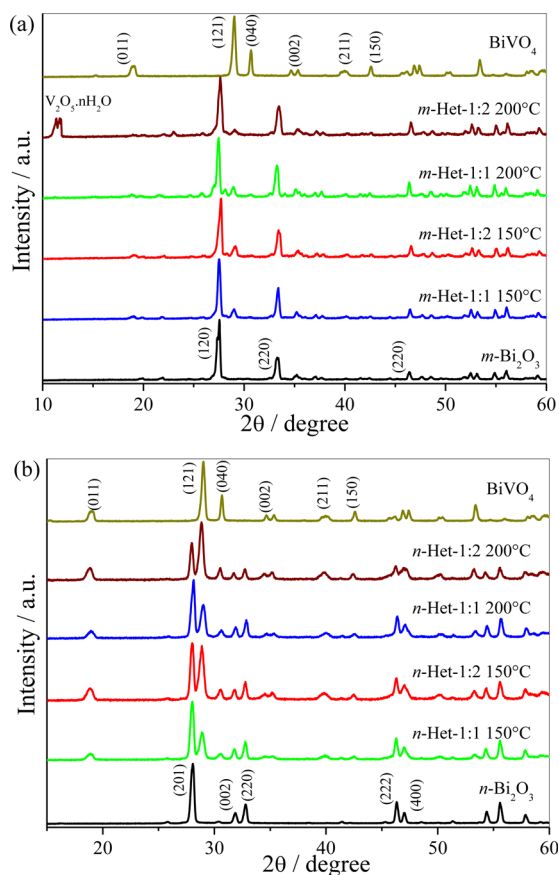


Figure 2. XRD patterns of (a) commercial $m\text{-Bi}_2\text{O}_3$, BiVO_4 , and the as-synthesized $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures and of (b) commercial $n\text{-Bi}_2\text{O}_3$, BiVO_4 , and the as-synthesized $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures.

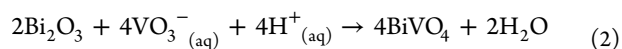
crystalline phases and to estimate the weight ratio between the phases. The hydrothermal treatment of $m\text{-Bi}_2\text{O}_3$ in the presence of the V precursor led to mixed crystalline phases that contained the monoclinic Bi_2O_3 (JCPDS, no. 41-1449) and monoclinic BiVO_4 (JCPDS, no. 83-1699). The weight percentage of BiVO_4 in each heterostructured sample was calculated using the relative intensities of the main peaks ($\% = I_{\text{BiVO}_4}(121)/(I_{\text{Bi}_2\text{O}_3}(120) + I_{\text{BiVO}_4}(121))$)⁴¹ identified in Figure 2a and Table 1. The increases in the concentration of the V precursor and the hydrothermal treatment temperature (except for the $m\text{-Het-1:2 200 }^\circ\text{C}$ sample) resulted in larger amounts of monoclinic BiVO_4 forming on the as-synthesized samples. The $m\text{-Het-1:2 200 }^\circ\text{C}$ sample showed a different behavior and formed a spurious phase (probably $\text{V}_2\text{O}_5 \cdot n\text{H}_2\text{O}$, as indicated by the peaks at $2\theta \approx 11.4^\circ$ and 11.6°) due to the excess of unreacted V precursor and the suitable thermal energy for the crystallization of vanadium compounds.^{42,43}

Table 1. Weight Percentage of the Monoclinic BiVO_4 Phase in the As-Synthesized Samples (Error < 3%)

samples	wt % BiVO_4^a	wt % BiVO_4^b
Het-1:1 150 $^\circ\text{C}$	16	33
Het-1:2 150 $^\circ\text{C}$	18	48
Het-1:1 200 $^\circ\text{C}$	18	37
Het-1:2 200 $^\circ\text{C}$	13	60

^aPercentage of the BiVO_4 phase in the heterostructure formed from $m\text{-Bi}_2\text{O}_3$. ^bPercentage of the BiVO_4 phase in the heterostructure formed from $n\text{-Bi}_2\text{O}_3$.

The effect of the Bi_2O_3 precursor particle size on heterostructure synthesis was evaluated using the nanometric Bi_2O_3 (Figure 2b). As observed earlier, despite the difference in the precursor particle size, a mixed-crystalline phase containing the tetragonal Bi_2O_3 (JCPDS, no. 27-0050) and monoclinic BiVO_4 (JCPDS, no. 83-1699) crystalline phases was obtained (Figure 2b). However, the use of the $n\text{-Bi}_2\text{O}_3$ precursor in heterostructure synthesis resulted in the formation of a larger amount of BiVO_4 compared to synthesis using the $m\text{-Bi}_2\text{O}_3$ precursor (Figure 2b, Table 1). The increases in the V precursor concentration and hydrothermal treatment temperature caused a larger amount of BiVO_4 to form on the heterostructure. Therefore, the growth of BiVO_4 on the Bi_2O_3 surface was proposed to follow a classical solubilization–precipitation reaction mechanism, as shown in eqs 1 and 2.



These results are in agreement with the findings of Liang et al.,⁴⁴ De-Kun et al.,⁴⁵ and Yu et al.,⁴⁶ who proposed synthesis routes driven by the solubility differences between Bi_2S_3 and $\text{Bi}_2\text{O}_2\text{CO}_3$, between Bi_2S_3 and BiVO_4 , and between CuS and ZnS , respectively. These studies used the principle that compounds with a high solubility can be converted into compounds with a low solubility, as observed by the formation of Bi_2S_3 from $\text{Bi}_2\text{O}_2\text{CO}_3$ and BiVO_4 due to its lower solubility compared to the other compounds.^{44,45,47} The crystallite size from the XRD data and the particle size from the specific surface area were calculated for the $m\text{-Bi}_2\text{O}_3$ and $n\text{-Bi}_2\text{O}_3$ samples (Table S1). The comparison between crystallite and particle sizes revealed that both samples are polycrystalline; therefore, it is expected that the particle size may be more influential on the solubility than the crystallite size. Samples $m\text{-Bi}_2\text{O}_3$ and $n\text{-Bi}_2\text{O}_3$ exhibited approximately the same crystallite size, which confirms that the expected differences in solubility are not related to the single-crystal solubility. We correlated Bi_2O_3 solubility to the particle size using the Ostwald–Freundlich equation

$$S = S_0 e^{(4\gamma_{\text{SL}}V_{\text{M}}/RTd)}$$

where S_0 is the solubility of the flat surface, γ_{SL} is the solid–liquid interfacial energy, V_{M} is the molar volume of the solid phase, R is the gas constant, d is the particle diameter, and T is the absolute temperature. The particle diameter (d) of the $m\text{-Bi}_2\text{O}_3$ and $n\text{-Bi}_2\text{O}_3$ samples was the only parameter that changed. Therefore, the higher solubility of $n\text{-Bi}_2\text{O}_3$ compared to $m\text{-Bi}_2\text{O}_3$ is related to the smaller particle size and higher surface area.

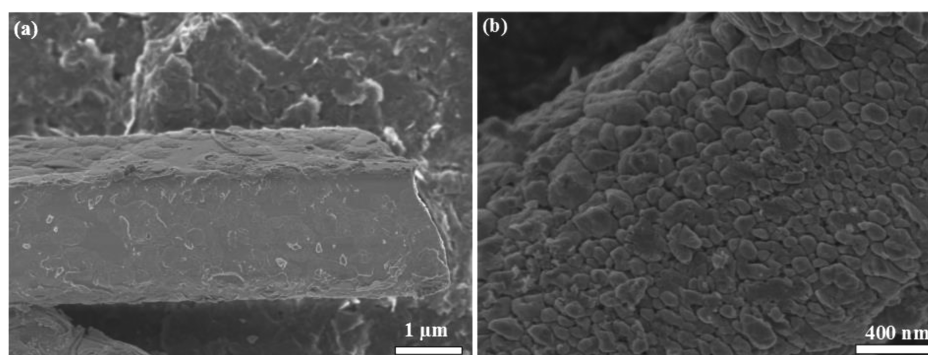


Figure 3. Representative FE-SEM images of (a) commercial $m\text{-Bi}_2\text{O}_3$ and (b) $m\text{-Het-1:1}$ 200 °C.

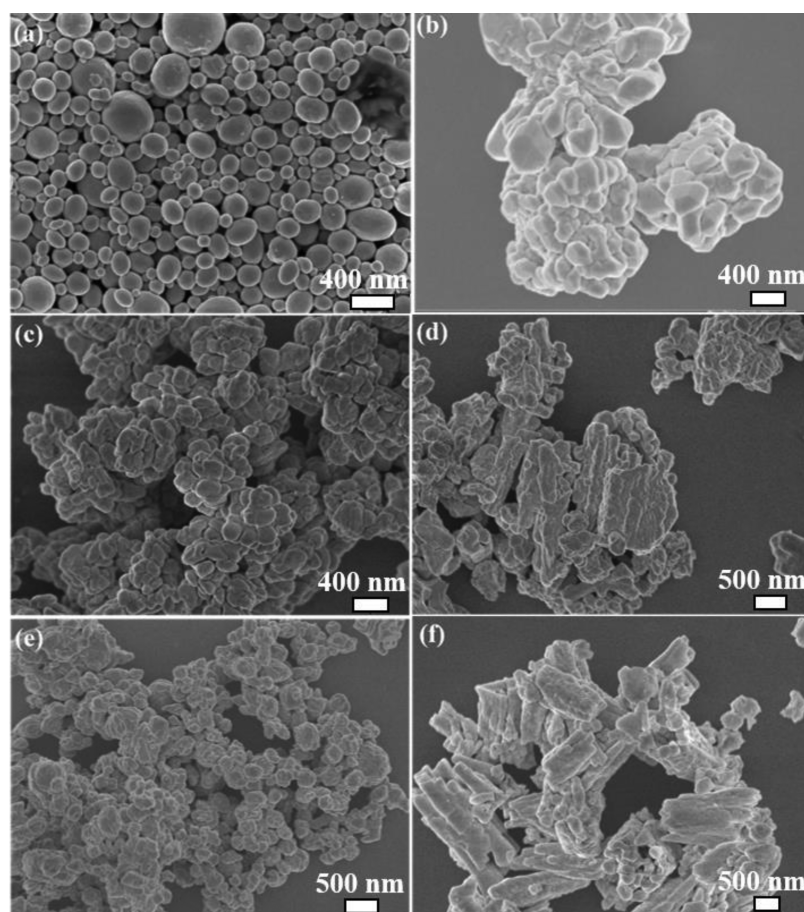


Figure 4. Representative FE-SEM images of (a) commercial $n\text{-Bi}_2\text{O}_3$, (b) BiVO_4 , (c) $n\text{-Het-1:1}$ 150 °C, (d) $n\text{-Het-1:2}$ 150 °C, (e) $n\text{-Het-1:1}$ 200 °C, and (f) $n\text{-Het-1:2}$ 200 °C.

Considering that NH_4VO_3 is highly water-soluble, adsorption to Bi_2O_3 is expected to be the first step in the surface reaction that occurs via partial dissolution and fast reprecipitation prior to the subsequent growth. However, this reaction must occur locally to ensure heterostructure formation because precipitation can lead to a mixture of phases. Representative SEM images of the as-synthesized heterostructured samples were analyzed to determine the morphology and confirm the crystallization/growth of the BiVO_4 nanoparticles on the m - and $n\text{-Bi}_2\text{O}_3$ surfaces (Figures 3 and 4, respectively). The $m\text{-Bi}_2\text{O}_3$ surface showed the presence of rodlike, micrometric particles with smooth surfaces. In the case of the $m\text{-Het-1:1}$ 200 °C sample, particles with a size and morphology similar to

those of the $m\text{-Bi}_2\text{O}_3$ precursor were observed, indicating that $m\text{-Bi}_2\text{O}_3$ was not fully solubilized and reprecipitated under this condition. However, the existence of a rougher surface composed of a large number of quasi-spherical nanoparticles could be attributed to the BiVO_4 phase. The EDX spectrum and elemental mapping image showed the presence of Bi and V in the $m\text{-Het-1:1}$ 200 °C sample (see the Supporting Information, Figure S1), as expected. Figure S1d shows that V was concentrated on the particle surface, confirming that the spherical nanoparticles were related to BiVO_4 and grew evenly on the $m\text{-Bi}_2\text{O}_3$ surface. In the case of the $m\text{-Het-1:2}$ 200 °C sample, the EDX spectrum and elemental mapping image of the Bi and V atoms (Figure S2) confirmed the formation of regions

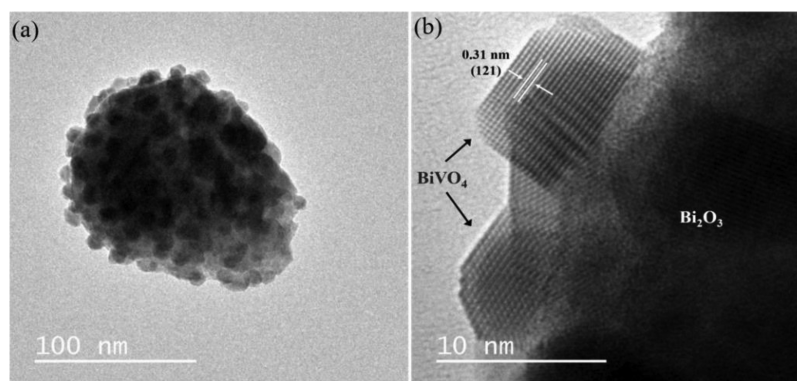


Figure 5. (a) TEM and (b) HRTEM images of the *n*-Het-1:1 200 °C sample.

with only V atom, which could be attributed to the spurious $V_2O_5 \cdot nH_2O$ crystalline phase identified by the XRD patterns. In addition, the growth of the phases was segregated, and the Bi and V atoms had different spatial positions, in agreement with the XRD observations.

The *n*- Bi_2O_3 sample exhibited a uniform, spherical morphology with an average particle size of 150 nm and a smooth surface,²⁵ as shown in Figure 4a. The *n*-Het-1:1 150 °C sample had particles with a size and morphology similar to those of *n*- Bi_2O_3 (Figure 4c), which confirmed that this oxide was not fully solubilized during the hydrothermal treatment at 150 °C. However, the surface of the *n*-Het-1:1 150 °C sample was rough, with a large number of nanospheres, which could be due to the formation of $BiVO_4$ on the *n*- Bi_2O_3 surface. The *n*-Het-1:2 150 °C sample had particles with a different morphology and larger particle size compared to those of the *n*- Bi_2O_3 precursor. This finding could be explained by considering that, under this synthesis condition, excess V precursor resulted in some *n*- Bi_2O_3 particles being completely converted to $BiVO_4$, and further growth led to phase segregation.

For the *n*-Het-1:1 200 °C sample, the particle size and morphology were very similar to those of the *n*- Bi_2O_3 precursor, but the sample had a rough surface, indicating that $BiVO_4$ grew on the Bi_2O_3 surface. This result was supported by the EDX spectrum and the elemental mapping image (Figure S3), which showed that, under this synthesis condition, the Bi and V atoms were homogeneously distributed in the particle and confirmed the growth of $BiVO_4$ on the Bi_2O_3 surface. The *n*-Het-1:2 200 °C sample presented a completely different morphology and particle size (Figure 4f) compared to those of the *n*- Bi_2O_3 precursor (Figure 4a), indicating that there was segregation of $BiVO_4$ after its formation, in agreement with the growth mechanism proposed for the *n*-Het-1:2 150 °C sample. This finding was confirmed by the EDX spectrum and the elemental mapping of the Bi and V atoms in the *n*-Het-1:2 150 °C heterostructured sample (Figure S4), where the V atoms were segregated in specific positions, suggesting that there were two different phases in the same particle. Therefore, excess V precursor during synthesis had a deleterious effect on the preparation of heterostructures via growth of $BiVO_4$ on the Bi_2O_3 surface because there was a high degree of segregation between the crystalline phases under these conditions; this effect was more pronounced at higher hydrothermal treatment temperatures.

Formation of the $Bi_2O_3/BiVO_4$ heterojunction was confirmed by the TEM and HRTEM images obtained for the *n*-

Het-1:1 200 °C sample (Figure 5). The TEM image revealed the presence of faceted nanoparticles with sizes smaller than 10 nm on the Bi_2O_3 surface. However, the HRTEM image of this sample (Figure 5b) showed that the nanoparticles of $BiVO_4$ and Bi_2O_3 coexisted in the same region in the monoclinic and tetragonal phases, respectively. The growth of $BiVO_4$ on the Bi_2O_3 surface was identified by its interlayer distance of 0.31 nm for the (121) plane. This result confirmed the formation of heterojunctions between Bi_2O_3 and $BiVO_4$.

Raman spectroscopy analyses were performed to determine the medium-range structures of the as-synthesized heterostructured samples. The Raman scattering spectra of the *m*- Bi_2O_3 precursor and the corresponding heterostructures are shown in Figure 6a. The *m*- Bi_2O_3 precursor showed the typical Raman spectrum of monoclinic, crystalline Bi_2O_3 with characteristic

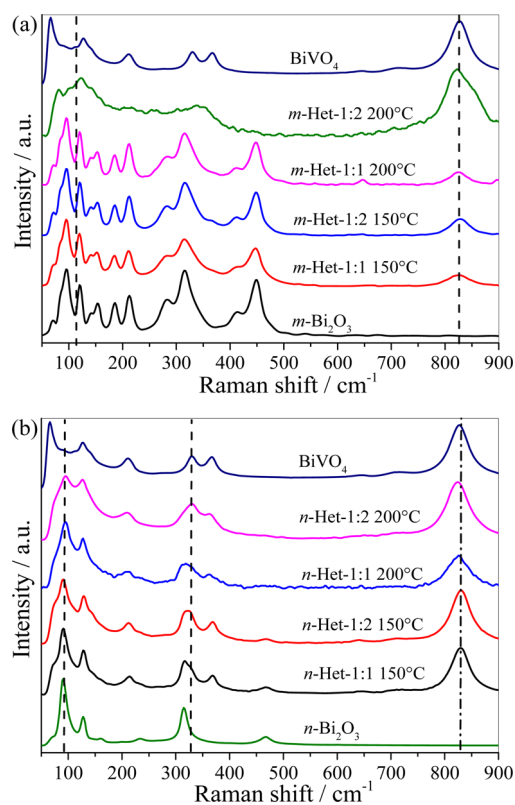


Figure 6. Raman scattering spectra of (a) *m*- Bi_2O_3 , $BiVO_4$, and the as-synthesized *m*- $Bi_2O_3/BiVO_4$ heterostructures and of (b) *n*- Bi_2O_3 , $BiVO_4$, and the as-synthesized *n*- $Bi_2O_3/BiVO_4$ heterostructures.

peaks at 70, 95, 122, 141, 153, 186, 213, 315, 413, 450, and 542 cm^{-1} , and these vibrational modes had A_g or B_g symmetry (or both).⁴⁸ The Raman scattering peaks observed in the region below 120 cm^{-1} were mainly assigned to shifts in the Bi atoms. The peaks in the range from 120 to 150 cm^{-1} were attributed to shifts in the Bi and O atoms, and those above 150 cm^{-1} were assigned to displacements of the O atoms.⁴⁹ The BiVO_4 sample showed a characteristic Raman spectrum with five main peaks at 120, 200, 330, 362, and 826 cm^{-1} .⁵⁰ The peaks at 120 and 200 cm^{-1} were assigned to the external vibrational mode of BiVO_4 , the peaks at 330 and 362 cm^{-1} were related to the asymmetric and symmetric deformation modes of the V–O bonds on tetrahedral VO_4 , respectively, and the peak at 826 cm^{-1} was assigned to the symmetric stretching mode of the V–O bonds with A_g symmetry.⁵¹ The Raman spectra of the heterostructured $m\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ samples showed peaks related to both the crystalline phases (Bi_2O_3 and BiVO_4), which was in good agreement with the XRD analysis. There was a gradual increase in the intensity of the peaks related to the BiVO_4 phase as the amount of the V precursor and the hydrothermal treatment temperature increased. The Raman spectrum of the $n\text{-Bi}_2\text{O}_3$ precursor showed six well-defined peaks at 70, 91, 127, 232, 315, and 467 cm^{-1} , which were attributed to tetragonal, crystalline Bi_2O_3 (Figure 6b), and the peaks at 70 and 91 cm^{-1} were related to the E_g and A_g symmetry vibration modes of the Bi atoms, respectively.^{48–50} The peaks at 127, 315, and 467 cm^{-1} were assigned to Bi–O bond stretching.^{48–50} In the case of the $n\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructured samples, the Raman spectra had peaks related to both of the crystalline phase constituents (Bi_2O_3 and BiVO_4), which was in good agreement with the XRD results.

Due to the fundamental role of the electronic properties of a semiconductor in heterogeneous photocatalysis applications, the band gap values of $m\text{-Bi}_2\text{O}_3$, $n\text{-Bi}_2\text{O}_3$, and their corresponding heterostructures were determined by applying the Tauc equation to the DRS data for the samples (Figure 7, Table 2). The $m\text{-Bi}_2\text{O}_3$ precursor had a band gap of 2.70 eV, which was in agreement with the values previously reported in the literature.⁵² The heterostructured $m\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ samples showed different band gap values with small fluctuations, and the values ranged from 2.35 to 2.40 eV. This result confirmed that BiVO_4 evenly decorated the surface of the $m\text{-Bi}_2\text{O}_3$ precursor because the band gap values are characteristic of monoclinic BiVO_4 . The results demonstrated that the Bi:V ratio and the hydrothermal treatment temperature did not affect the band gap characteristics. The band gap of the $n\text{-Bi}_2\text{O}_3$ precursor was 2.60 eV (Table 2).⁵² Different band gap features were observed for the $n\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ samples (Figure 7b). The heterostructured $n\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ samples showed two band gaps related to BiVO_4 and $n\text{-Bi}_2\text{O}_3$, with values of approximately 2.40 and 2.60 eV, respectively.

As expected, the specific surface areas of the as-synthesized samples (Table 3) decreased due to particle growth upon the formation of BiVO_4 on the Bi_2O_3 surface, which was observed in the SEM and TEM images (Figures 3–5). This finding indicated that any positive effect on the photocatalytic performance of the heterostructures was not related to the specific surface area.

The photocatalytic properties of the $m\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures were evaluated using the photodegradation of MB dye, and the results were compared to those obtained using the pure $m\text{-Bi}_2\text{O}_3$ sample (Figure 8a). All the $m\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures were photoactive for MB dye photo-

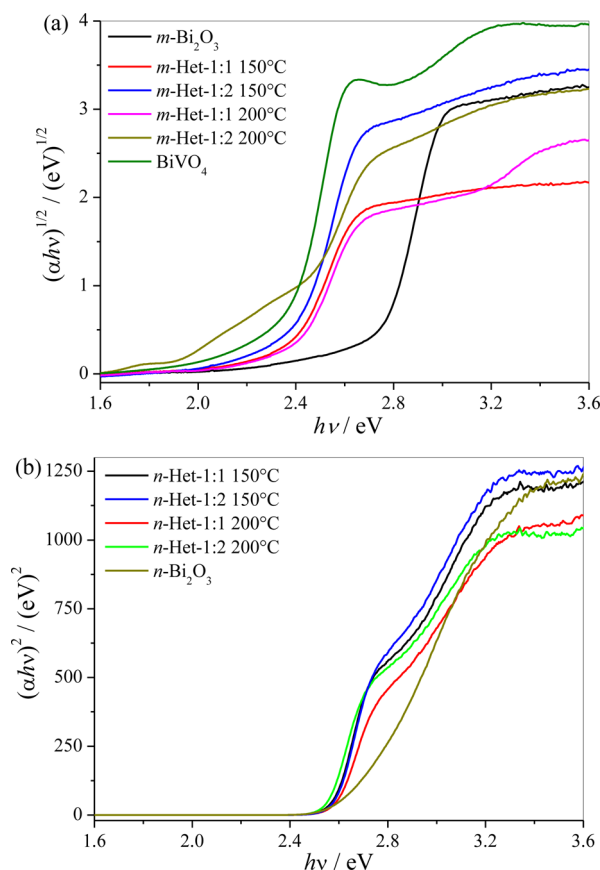


Figure 7. Band gap determination using indirect and direct Tauc plots with the DRS data for (a) $m\text{-Bi}_2\text{O}_3$, (b) $n\text{-Bi}_2\text{O}_3$, and the corresponding heterostructured samples.

Table 2. Band Gap Values of Pure $m\text{-Bi}_2\text{O}_3$ and $n\text{-Bi}_2\text{O}_3$ and Their Corresponding Heterostructures

samples	band gap (eV) ^a	band gap (eV) ^b
Bi_2O_3	2.70 ± 0.09	2.60 ± 0.01
Het-1:1-150	2.35 ± 0.04	2.50 ± 0.03
Het-1:2-150	2.40 ± 0.10	2.40 ± 0.01
Het-1:1-200	2.35 ± 0.04	2.50 ± 0.02
Het-1:2-200	2.30 ± 0.03	2.40 ± 0.02

^aBand gap values of the heterostructures formed from $m\text{-Bi}_2\text{O}_3$, ^bBand gap values of the heterostructures formed from $n\text{-Bi}_2\text{O}_3$

Table 3. Specific Surface Area Values (SSAs) of the $m\text{-Bi}_2\text{O}_3$ and $n\text{-Bi}_2\text{O}_3$ Precursors and Their Most Photoactive Heterostructures

samples	SSA ($\text{m}^2\cdot\text{g}^{-1}$)
$m\text{-Bi}_2\text{O}_3$	0.15
$m\text{-Het-1:1 200 }^\circ\text{C}$	0.10
$n\text{-Bi}_2\text{O}_3$	2.80
$n\text{-Het-1:1 200 }^\circ\text{C}$	0.40
BiVO_4	0.60

degradation and showed higher activity than the $m\text{-Bi}_2\text{O}_3$ precursor and the blank experiment (both conditions exhibited negligible photoactivity). The blank experiment in our study was an aqueous solution of the MB dye without any photocatalyst. On the basis of the reaction rate constants, which were calculated using the pseudo-first-order equation

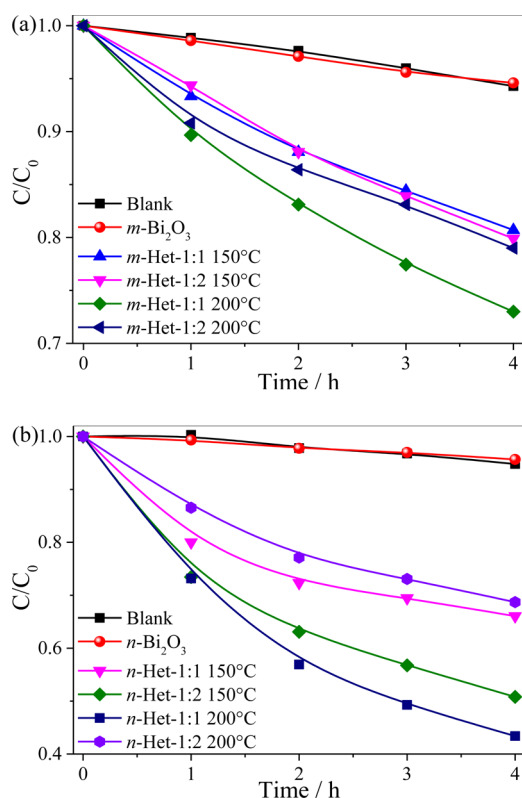


Figure 8. Kinetics curves for the MB dye ($5 \text{ mg}\cdot\text{L}^{-1}$) photodegradation catalyzed by (a) $m\text{-Bi}_2\text{O}_3$ and the $m\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures and by (b) $n\text{-Bi}_2\text{O}_3$ and the $n\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures.

(Table 4), the photoactivity followed the order: $m\text{-Het-1:1 } 200^\circ\text{C} > m\text{-Het-1:2 } 200^\circ\text{C} \approx m\text{-Het-1:2 } 150^\circ\text{C} \approx m\text{-Het-1:1 } 150^\circ\text{C}$

Table 4. First-Order Rate Constants ($k \times 10^2$) for the MB Dye Photodegradation Reaction under Visible Irradiation Catalyzed by Bi_2O_3 and the Heterostructures

sample	$m\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ $k \times 10^2 \text{ (h}^{-1}\text{)}$	$n\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ $k \times 10^2 \text{ (h}^{-1}\text{)}$
Bi_2O_3	1.3 ± 0.2	1.3 ± 0.2
Het-1:1 150°C	5.0 ± 0.3	10.5 ± 0.5
Het-1:2 150°C	5.4 ± 0.3	18.0 ± 2.0
Het-1:1 200°C	7.1 ± 0.7	22.5 ± 1.9
Het-1:2 200°C	5.3 ± 0.4	11.0 ± 2.0

$^\circ\text{C}$. The heterostructure with the largest amount of BiVO_4 showed the highest photoactivity for MB dye degradation, indicating that effective heterojunctions formed between the $m\text{-Bi}_2\text{O}_3$ and BiVO_4 phases increased the charge carrier lifetime. Figure 8b shows the MB dye photodegradation catalyzed by the $n\text{-Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures. As previously observed for the $m\text{-Het}$ heterostructures, all the heterostructures showed higher photoactivity than the precursor ($n\text{-Bi}_2\text{O}_3$). On the basis of the reaction rate constants (Table 4), the photoactivity of the heterostructures was in the following order: $n\text{-Het-1:1 } 200^\circ\text{C} > n\text{-Het-1:2 } 150^\circ\text{C} > n\text{-Het-1:1 } 150^\circ\text{C} \approx n\text{-Het-1:2 } 200^\circ\text{C}$. For these samples, an intermediate composition between the $n\text{-Bi}_2\text{O}_3$ and BiVO_4 phases resulted in improved photoactivity. The higher activity of the $n\text{-Het-1:1 } 200^\circ\text{C}$ sample was due to an ideal amount of BiVO_4 (37 wt %) and better distribution of BiVO_4 on the Bi_2O_3 surface because the heterostructures

obtained using excess V precursor ($n\text{-Het-1:2 } 150^\circ\text{C}$ and $n\text{-Het-1:2 } 200^\circ\text{C}$) had high amounts of BiVO_4 and segregation between the phases, as observed in the SEM images and the EDX elemental mapping.

Despite the obvious formation of heterojunctions between the Bi_2O_3 and BiVO_4 phases, an additional investigation comparing the heterostructures ($m\text{-Het-1:1 } 200^\circ\text{C}$ and $n\text{-Het-1:1 } 200^\circ\text{C}$) with the two isolated phases (Bi_2O_3 and BiVO_4) and the corresponding physical mixtures was needed to confirm the effect of the heterojunctions and formation of the type-II heterostructure. The $m\text{-Het-1:1 } 200^\circ\text{C}$ heterostructure exhibited lower activity for MB dye photodegradation compared to that of pure BiVO_4 (Figure 9a). However, the

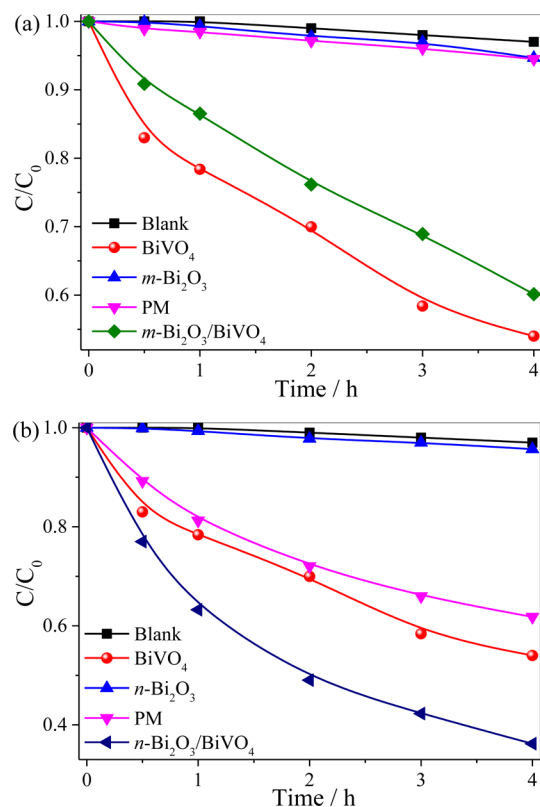


Figure 9. Photocatalytic performance in MB dye ($5 \text{ mg}\cdot\text{L}^{-1}$) degradation using the heterostructures obtained with the precursors (a) $m\text{-Bi}_2\text{O}_3$ and (b) $n\text{-Bi}_2\text{O}_3$ as well as their corresponding physical mixtures (PMs) and the isolated phases.

heterostructure showed higher photoactivity than the physical mixture of BiVO_4 and $m\text{-Bi}_2\text{O}_3$ prepared in the same proportions as $m\text{-Het-1:1 } 200^\circ\text{C}$. This heterostructured sample showed a lower activity than pure BiVO_4 due to the large amount of the $m\text{-Bi}_2\text{O}_3$ precursor (82 wt %), which exhibits insignificant photoactivity for MB dye photodegradation under visible irradiation. The $n\text{-Het-1:1 } 200^\circ\text{C}$ sample showed superior photoactivity for MB dye degradation (Figure 9b) compared to the isolated phases ($n\text{-Bi}_2\text{O}_3$ and BiVO_4) and the corresponding physical mixture. Hence, the as-synthesized $m\text{-Het-1:1 } 200^\circ\text{C}$ and $n\text{-Het-1:1 } 200^\circ\text{C}$ samples both showed higher photocatalytic performances than their corresponding physical mixtures, indicating that the formation of the heterojunctions increased the charge carrier lifetimes probably due to type-II heterostructure formation.

The results showed that the heterostructures obtained using the n - Bi_2O_3 precursor were more efficient than the heterostructures obtained using the m - Bi_2O_3 precursor. The n - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures provided an MB dye degradation rate approximately 2-fold higher than that obtained using the m - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures. This result was expected based on the previous analysis because the heterostructures that were synthesized using the nanometric Bi_2O_3 particles possessed a larger amount of BiVO_4 on their surfaces than the heterostructures synthesized using the micrometric Bi_2O_3 particles. This difference resulted in the increased formation of heterojunctions between the Bi_2O_3 and BiVO_4 phases. These effects on photocatalytic performance can be clearly seen from the pseudo-first-order reaction rate constants for MB photodegradation as a function of the amount of the BiVO_4 phase in the heterostructures (Figure 10).

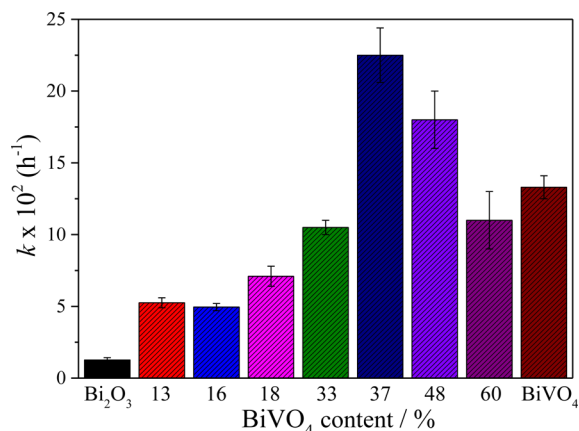


Figure 10. Pseudo-first-order rate constants for the MB dye photodegradation reaction as a function of the BiVO_4 content in the heterostructures.

Finally, to confirm the formation of the n - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructure, the electron/hole pair lifetime was calculated using time-resolved photoluminescence (TRPL). The photoluminescence decay spectra at 545 nm for the BiVO_4 and n - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ samples are shown in Figure 11. The TRPL

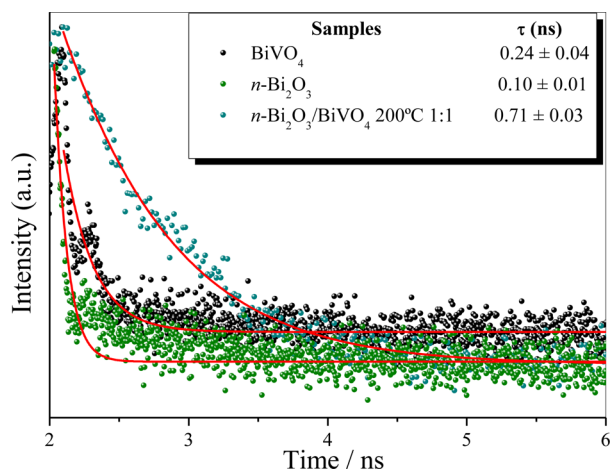


Figure 11. Observed (points) and calculated (lines) time-resolved photoluminescence intensities (decay lifetimes) for the n - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$, BiVO_4 , and n - Bi_2O_3 samples. The powders were excited at 405 nm, and photoluminescence was monitored at 545 nm.

decays for both samples were fitted using a first-order exponential function to calculate the charge carrier lifetime. The electron/hole pair lifetime of the n - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructure (Het-1:1 200 °C) was significantly higher than that of the pure BiVO_4 and Bi_2O_3 . The lifetime of the band–band emission (the electron/hole pair recombination at 545 nm) of the n - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructure was 0.71 ns, while the lifetimes for BiVO_4 and Bi_2O_3 were approximately 0.24 and 0.10 ns, respectively. These results showed that the lifetime of the electron/hole pair in the n - $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructure was 2.9-fold and 7.1-fold longer than that in pure BiVO_4 and Bi_2O_3 , respectively, proving that the junction of Bi_2O_3 with BiVO_4 formed a suitable type-II heterostructure.

The conduction band (CB) and valence band (VB) potentials of BiVO_4 are 0.3 and 2.7 eV, respectively.^{38,53} The CB and VB potentials of tetragonal Bi_2O_3 are 2.9 and 0.6 eV, respectively.^{53,54} BiVO_4 is an intrinsic n-type semiconductor, while Bi_2O_3 is as an intrinsic p-type semiconductor.^{26,55} Thus, before the semiconductors come into contact, the conduction and valence band edges of the p-type Bi_2O_3 are lower than those of the n-type BiVO_4 , and the Fermi level of the Bi_2O_3 is lower than that of the BiVO_4 (Figure S5). After the semiconductors come into contact, the Fermi level of Bi_2O_3 moves up, while the Fermi level of BiVO_4 moves down until the Fermi levels of Bi_2O_3 and BiVO_4 become the same.^{37,32,33} A type-II heterojunction (p–n junction) is formed, and electron transfer occurs from Bi_2O_3 to BiVO_4 until their Fermi levels align; i.e., the semiconductor system reaches a thermal equilibrium state. As a result, the whole energy band of Bi_2O_3 increases while that of BiVO_4 decreases. In the end, Bi_2O_3 has a conduction band and valence band higher than BiVO_4 . Therefore, the interface between the phases allows transfer of the photogenerated electrons from the p-type (Bi_2O_3) to the n-type (BiVO_4) semiconductor. The photogenerated holes are transferred from the n-type (BiVO_4) to the p-type (Bi_2O_3) semiconductor because the valence band level of BiVO_4 is lower than that of Bi_2O_3 .^{32–34} Hence, the photogenerated electrons accumulate in the BiVO_4 conduction band, while the photogenerated holes accumulate in the Bi_2O_3 valence band, which is where the MB dye oxidation reaction occurs. The enhanced photocatalytic activity of the $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures can be explained by the spatial separation of the photogenerated electron/hole pair.

4. CONCLUSIONS

An alternative method based on solubility differences was developed using preformed micro- or nanometric particles of Bi_2O_3 to obtain $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures. The technique enabled efficient heterojunction tailoring, and the amount of BiVO_4 in the heterostructure was dependent on the particle size of the preformed Bi_2O_3 due to the solubility differences between the micro- and nanometric Bi_2O_3 . The growth mechanism of BiVO_4 on the self-sacrificial Bi_2O_3 surface was elucidated, and the formation of the $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterojunctions was confirmed using HRTEM images. The photocatalytic activity was enhanced compared to the isolated phases due to the formation of the $\text{Bi}_2\text{O}_3/\text{BiVO}_4$ heterostructures. The formation of the type-II heterostructure increased the spatial separation of the photogenerated electron/hole pair, which increased its lifetime, as demonstrated by the PL and TRPL analyses.

■ ASSOCIATED CONTENT**Supporting Information**

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SEM images, EDX spectra, and elemental mapping (PDF)

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The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

Notes

The authors declare no competing financial interest.

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